

INTERCONNECTING CHANNEL MORPHOLOGY  
COMPOSITION FOR RELEASING CO<sub>2</sub>

Inventor: Ihab Hekal  
230 Round Hill Road  
Greenwich, Connecticut 06831

Attorney Docket No.: 5094.053

1003397-121901  
FOOTER

## **RELATED APPLICATIONS**

The present application claims priority to U.S. Serial No. 60/256,677, entitled Interconnecting Channel Morphology Composition For Releasing CO<sub>2</sub>, filed December 19, 2000.

## **FIELD OF THE INVENTION**

The present invention is directed to compositions, which in a period of time release carbon dioxide (CO<sub>2</sub>) gas. The invention may be used in caps and bottles used to contain carbonated beverages. Such bottles are known to not be impervious, and leak CO<sub>2</sub> over a period of time. In one embodiment, the composition has a co-continuous interconnecting channel morphology comprising three components - - two polymers (i.e. components A and B) and a particle (i.e. component C) wherein the channels consist mainly of component B and the majority of component C resides in the channels.

## **Background of the Invention**

Carbonated soft drinks (CSD) bottles are often packaged in bottles constructed of PET. A non-exhaustive list of such beverages includes colas, seltzers, club soda, ginger ale, Tom Collins mix, root beers, lemon-lime flavored sodas, and a variety of other flavored sodas. At pressures near or below atmosphere, the CO<sub>2</sub> in these carbonated beverages escapes rapidly from solution. To keep a sufficient amount of CO<sub>2</sub> in solution, the soda is bottled under pressure. Keeping the CO<sub>2</sub> in solution is relatively important. Without it, the soda becomes flat, and these beverages would lose much of their appeal. Anecdotal evidence suggests that it is the carbonation that makes the beverages appealing. That is, the CO<sub>2</sub> present in the beverage provides a pleasing sensation in the palate. Accordingly, there is a need to insure that a sufficient amount of CO<sub>2</sub> remains in the beverage so that the consumer realizes full value for the product.

PET is the material of choice for forming bottles for carbonated beverages, due to its low cost, transparency, moldability, and relatively light weight. Also, its high level of impact resistance renders PET containers virtually unbreakable, a desired feature in a packaging material. However, the barrier properties of PET with respect to gases are somewhat suspect, and over time, CO<sub>2</sub> will leak through the container. Also, such containers are typically sealed by a cap having a base and a cylindrical skirt extending downward from the base. The cap has threads on the interior of the skirt, which are complimentary to threads on the container located near the container opening.

It is possible for the pressurized CO<sub>2</sub> to escape from the container in the neck region as well. Since it is possible that a beverage may sit in storage for an extended period of weeks, even

months, between bottling and use by the consumer, a significant loss of CO<sub>2</sub> may arise. In one example, the shelf life of a 12 fluid ounce bottle is 8 weeks and 20 fluid ounce bottle is 16 weeks.

Furthermore, once the consumer opens the bottle, the CO<sub>2</sub> is rapidly released. In the case of large sized bottles, which hold a volume in excess of a single serving, it is expected that these bottles will be opened, closed after beverage is dispensed, and then opened at a later time to again dispense beverage. Eventually all of the CO<sub>2</sub> in the beverage will be released from the beverage rendering the remaining soda fizzless, a wholly insipid and undesirable state. The consumer is then left with two options, neither satisfactory, in which he or she can consume a flat beverage, or dump the remainder down the drain.

### **DESCRIPTION OF THE DRAWINGS**

Fig. 1 is a graph that illustrates the results for CO<sub>2</sub> release for Example 1 and 2.

Fig. 2 is a graph that illustrates the results for the leakage test for the control samples for Examples 1 and 2.

Fig. 3 is a graph that illustrates the results for Example 3.

### **SUMMARY OF THE INVENTION**

The present invention is a CO<sub>2</sub> releasing composition that can be incorporated into the packaging for carbonated beverages. In one embodiment, the composition contains a CO<sub>2</sub> releasing component. In another embodiment the composition is formed into a film. In yet another embodiment, the present invention is three phase moisture-activated polymer film that releases CO<sub>2</sub>. In a further embodiment, this film can be incorporated into the inner seals of soda bottle cap easily and safely, so that in the presence of the water in the beverage the film releases CO<sub>2</sub>. With this arrangement, there is always certain amount of carbon dioxide present inside the soda bottle due to the carbon dioxide releasing characteristics of the film, which can enhance the shelf life of the bottle significantly.

In another embodiment, the composition has a co-continuous interconnecting channel morphology comprising three components - - two polymers (i.e. components A and B) and a particle (i.e. component C) wherein the channels consist mainly of component B and the majority of component C resides in the channels. Components A and B are generally immiscible within each other. In addition, one criteria for selecting component C and components A and B may be based on component C preferential affinity for component B over component A. Another criteria for selecting component C is based on the capacity of component C to release CO<sub>2</sub>. In yet another embodiment, the CO<sub>2</sub> releasing component releases CO<sub>2</sub> when it come in contact with moisture, such as water, including water vapor.

## **DETAILED DESCRIPTION OF THE INVENTION**

The releasing composition of the present invention is generally incorporated into a plastic. Therefore, in one embodiment, the releasing composition can be applied to the liner that is often placed on the bottom of the base of the cap, which faces into the bottle when the cap is fixed to the top of the bottle. The composition can be die cut to fit within the interior of the cap. The composition can be affixed to the interior liner of the cap by any known means, such as heat sealing, or by use of adhesives. In another embodiment, the liner may be coextruded. In a further embodiment, the composition can be formed into an induction seal, such as disclosed in U.S. Provisional Application 60/225,713, which is incorporated by reference herein. The composition can also be applied to the interior surface of the bottle. However, this may not be desirable, since the inclusion of the composition will render those locations nontransparent the region where the material is located.

In one embodiment, the composition comprises a co-continuous interconnecting channel morphology comprising at least three components, wherein: (a) component A is selected from the group of polymers that are semicrystalline polymers and amorphous polymers, wherein the amorphous polymers have a shear modulus greater than about 8 MPa; (b) component B is a polymer; (c) components A and B are immiscible within each other, and if components A and B react after mixing, components A and B are immiscible prior to reacting; (d) component C is a particle; (e) the volume fraction of component A represents at least about 50% by volume of the total volume of components A, B and C; (f) the preferential affinity between component B and component C is greater than between component A and component C; (g) at least two phases are formed, one phase is composed of a majority of component A, and the second phase is composed of a majority of components B and a majority of component C; and (h) two phases form the co-continuous interconnecting channel morphology.

Components A, B and C may be selected based on the desired end-use result -- the desired end-use property. For example, component A may typically be selected based on its permeability properties (e.g. barrier properties), its chemical and/or temperature resistance properties, its molding properties, and/or its price (e.g. since it is the component having the largest volume fraction of the composition). Similarly, for example, component B may typically be selected based on its transport properties (e.g. desired transport of CO<sub>2</sub>) and/or its preferential affinity with component C. Also, for example, component C is based on its ability to release CO<sub>2</sub>. Consequently, a specific composition may be uniquely tailored and thus, uniquely optimized for a desired end-use application.

For example, one method of forming the composition of the present invention is by adding component C and component B to component A when component A is in a molten state; or before component A is in the molten state, so that components B and C may be blended and thoroughly mixed throughout component A to insure that the blend is uniformly mixed before reaching the melt phase. For example, such a technique may be useful when components A, B and C are all powders. In another embodiment, component B and component A are mixed prior to adding component C. Component B is either added before component A is in the molten state or after component A is in the molten state. For example, component C may be added to component A during the thermal process of forming sheets. After blending and processing, component B is drawn out into interconnecting channels that contain a percolation path in component A. The majority of component C resides in the interconnecting channels because of its preferential affinity towards component B over component A. In addition, the composition of the present invention may be described as "monolithic" because the composition does not consist of two or more discrete macroscopic layers.

For purposes of the present invention, the term "phase" means a portion of a physical system that is uniform throughout, has defined boundaries and, in principle, can be separated physically from other phases. The term "melting point" is defined as the first order transition point of the material determined by DSC. The term "not mutually soluble" means immiscible with each other. The term "immiscibility" means that the components of the blend are driven by thermodynamic forces to separate (*i.e.* demix) into two or more distinct phases that will coexist indefinitely under equilibrium conditions. An example is the separation of the oil-rich and water-rich phases in a salad dressing. For purposes of the present invention, "partial" immiscibility or "partial" miscibility is deemed "immiscible" and thus, any tendency for a component to phase separate from another component is deemed "immiscible." Immiscibility may be determined by the application of one or more forms of microscopy (e.g., optical, TEM, SEM or AFM) with an observation that the components are separated into two or more distinct phases. The term "particle" means a dispersed component that is either a crystalline or amorphous solid, or a crosslinked organic or inorganic material, and that retains its shape, aside from recoverable deformations, before, during, and after the blend is compounded in the molten state at elevated temperatures. This would include, e.g., a crosslinked polymer latex.

Further, for purposes of the present invention, the term "co-continuous interconnecting channel morphology" means that the minor phase (*i.e.*, component B) is drawn out into interconnected channels that contain a percolation path, while simultaneously, the majority phase

(i.e., component A) is percolating. "Percolation" means that there exists at least one unbroken path, composed only of points from within that phase, that will lead from any surface of a sample through the interior of the sample to any other surface. Such a percolation path provides a route for a desired object, such as a small molecule, an atom, an ion, or an electron, to be macroscopically transported across the sample while contacting only one of the phases. For some systems, the existence of an interconnecting channel morphology that is co-continuous may be determined by a minimum of two transport measurements that demonstrate percolation in both minor and major phases. Percolation theory is a mature branch of mathematics and physical science that is described in a variety of review articles, specialized monographs, and many introductory texts on stochastic processes, probability theory, and statistical mechanics. For example, an introductory treatment of percolation theory is described by *D. Stauffer* in *Introduction to Percolation Theory*, Taylor and Francis, (London 1985).

The term "preferential affinity" means that the particle (i.e., component C) has a lower interfacial energy when contacting one component than compared to contacting another component. A suitable method for determining "preferential affinity" for the present invention is the following:

- (a) Blend the particle with the two components at elevated temperatures in their liquid state. Mix to achieve a macroscopically homogeneous dispersion.
- (b) Cool the mixture and allow to solidify.
- (c) Use a form of microscopy (e.g., TEM, SEM, and/or AFM) on a thin section to determine which of the two phases most closely contacts each particle in the field of view.
- (d) The component that is in the majority in the phase that contacts the largest number of particles is the component with "preferential affinity" for the particle.

Further, the term "shear modulus" is the ratio of a measured shear stress to the magnitude of a small, elastically recoverable, shear strain that is used to produce that stress. The criterion of greater than about 8 MPa refers to the shear modulus measured at room temperature. The "shear modulus" is determined by ASTM test method E143-87 (1998). The term "polymer" means a composition that is made by reacting two or more molecular species ("monomers") to form chemically-bonded larger molecules. The term "semicrystalline" means that the polymeric component, at ambient temperature, contains regions in which chain segments are packed with spatial registry into a periodic lattice and these regions are of sufficient size and extent to exhibit a detectable melting endotherm in a differential scanning calorimetry (DSC) measurement. The

term "amorphous" means that the polymeric component, at ambient temperature, either contains no regions of periodic packing of segments, or such regions are undetectable with a DSC measurement.

With respect to component A, in one embodiment, component A may be a thermoplastic material. Examples of suitable thermoplastic materials may include polyolefins such as polypropylene and polyethylene, polyisoprene, polybutadiene, polybutene, polysiloxane, polycarbonates, polyamides, ethylene-vinyl acetate copolymers, ethylene-methacrylate copolymer, poly(vinyl chloride), polystyrene, polyesters, polyanhydrides, polyacrylonitrile, polysulfones, polyacrylic ester, acrylic, polyurethane and polyacetal, or copolymers or mixtures thereof.

In one embodiment, component "B" may be either an oxide such as ethylene oxide or a polymer such as acrylic or a combination thereof. The loading of component "B" may range from about 10% to 20%, 20% to 40% and 40% to 60% by weight with respect to the polymer.

In yet a further embodiment, the CO<sub>2</sub> releasing component may be a carbon and oxygen containing salt of an alkali or alkaline earth metal, such as the carbonates of sodium, potassium, magnesium, or calcium. Bicarbonates of the same may also be employed. In yet a further embodiment, the CO<sub>2</sub> releasing component is calcium carbonate. In yet a further embodiment, the CO<sub>2</sub> releasing component may be baking powder. Mixtures of CO<sub>2</sub> releasing components are also possible. In an embodiment, the CO<sub>2</sub> releasing material loading level can range from about 10% to 20%, 20% to 40% and 40% to 60% by weight with respect to the polymer.

In another embodiment, the CO<sub>2</sub> releasing component may be a combination of a carbonate such as, but not limited to, sodium bicarbonate and/or calcium carbonate with an acid. Suitable acids include, but are not limited to, citric acid, maleic acid, malic acid, fumaric acid, polyacrylic acid, oxalic acid and/or mixtures thereof. In one specific embodiment, the acid can be used in an anhydrous form to prevent the reaction from occurring prematurely. In yet another embodiment, to reduce the possibility of the reaction initiating prematurely, the acid is in its solid granular form and is substantially free of moisture. In yet a further embodiment, the CO<sub>2</sub> releasing component may be combined with the plastic component without the employing the co-continuous interconnecting channel morphology comprising at least three components discussed above. In one example, this can be used when one desires to slow down the CO<sub>2</sub> release.

In a further embodiment, the acid particles may be employed as component "C" in creating co-continuous interconnecting channel morphology comprising at least three

components. Consequently, in this embodiment, the composition should be processed at a temperature below the melt temperature of the acid so that the acid remains in its particle form.

In yet another embodiment, the acid can be selected with a low molecular weight and thus, will allow multiple H<sup>+</sup> ions to be released per mole of the compound. In this way, less acid by weight may be used to achieve substantial completion of the chemical reaction. It is also understood that particle size of the acid and/or carbonate can affect the efficiency of the reaction. For example, a particle size in the range of about 1 micron for calcium carbonate can be employed.

The following are merely illustrative examples of various compositions of the present invention. It is understood that these are merely illustrative examples and are not meant to limit the present invention.

		Materials	Wt percentage
		CaCO <sub>3</sub>	40.00%
		PolyOx 750	3.00%
		Carbopol 934 p-nf	30.00%
		EXACT 4023	27.00%
		CaCO <sub>3</sub>	40.00%
		PolyOx 750	3.00%
		Carbopol 934 p-nf	20.00%
		EXACT 4023	37.00%
		B1H4E Resin (Hekal)	95.00%
		PolyOx 750	5.00%
		CaCO <sub>3</sub>	50.00%
		Luviskol VA64	7.00%
		Carbopol 971 P-NF	13.00%
		EXACT 4023	30.00%
		CaCO <sub>3</sub>	40.00%
		Luviskol VA64	10.00%
		Molecular Sieve 4A	15.00%
		EXACT 4023	35.00%
		NaHCO <sub>3</sub>	35.00%
		PolyOx 750	5.00%
		Carbopol 971 P-NF	25.00%
		EXACT 4023	35.00%
		NaHCO <sub>3</sub>	32.00%
		Citric Acid	28.00%
		PolyOx 750	5.00%
		EXACT 4023	35.00%
		CaCO <sub>3</sub>	40.00%



		PolyOx 750	5.00%
		Carbopol 971 P-NF	25.00%
		EXACT 4023	30.00%
		NaHCO <sub>3</sub>	37.00%
		Malic Acid	37.00%
		PolyOx 750	5.00%
		EXACT 4023	21.00%
		NaHCO <sub>3</sub>	50.00%
		Malic Acid	25.00%
		PolyOx 750	5.00%
		EXACT 4023	20.00%
		NaHCO <sub>3</sub>	37.00%
		Malic Acid	37.00%
		PolyOx 750	5.00%
		FT120WV3	21.00%
		NaHCO <sub>3</sub>	37.00%
		Malic Acid	37.00%
		PolyOx 750	5.00%
		TI 4040G	21.00%
		NaHCO <sub>3</sub>	37.00%
		Malic Acid	37.00%
		PolyOx 750	5.00%
		Dowlex 2045A	21.00%
		NaHCO <sub>3</sub>	37.00%
		Malic Acid	37.00%
		PolyOx 750	5.00%
		Elvax 3185	21.00%
		NaHCO <sub>3</sub>	37.00%
		Malic Acid	37.00%
		PolyOx 750	5.00%
		4006	21.00%
		NaHCO <sub>3</sub>	37.00%
		Malic Acid	37.00%
		PolyOx 750	5.00%
		Engage 4006	21.00%
		NaHCO <sub>3</sub>	37.00%
		Malic Acid	37.00%
		15-200	5.00%
		EXACT 4023	21.00%
		NaHCO <sub>3</sub>	37.00%
		Malic Acid	37.00%
		Luviskol Va64	5.00%
		EXACT 4023	21.00%
		NaHCO <sub>3</sub>	50.00%
		PolyOx 750	7.00%

		EXACT 4023	43.00%
		NaHCO <sub>3</sub>	45.00%
		Silica Grade 11	15.00%
		PolyOx 750	7.00%
		EXACT 4023	33.00%
		NaHCO <sub>3</sub>	37.00%
		Malic Acid	37.00%
		Pluronic F108	5.00%
		EXACT 4023	21.00%
		NaHCO <sub>3</sub>	45.00%
		Silica Grade 11	15.00%
		Pluronic F108	5.00%
		EXACT 4023	35.00%
		NaHCO <sub>3</sub>	40.00%
		Silica Grade 11	30.00%
		PolyOx 750	5.00%
		EXACT 4023	25.00%
		Calcium Carbonate	40.00%
		Silica Grade 11	30.00%
		PolyOx 750	5.00%
		EXACT 4023	25.00%
		Calcium Carbonate	45.00%
		Silica Grade 11	15.00%
		Pluronic F108	5.00%
		EXACT 4023	35.00%
		Calcium Carbonate	37.00%
		Citric Acid	31.00%
		Ca(OH) <sub>2</sub>	6.00%
		PolyOx 750	5.00%
		EXACT 4023	21.00%
		Calcium Carbonate	37.00%
		Malic Acid	37.00%
		Pluronic F108	5.00%
		EXACT 4023	21.00%
		NaHCO <sub>3</sub>	35.00%
		Malic Acid	35.00%
		PolyOx 750	10.00%
		EXACT 4023	20.00%
		NaHCO <sub>3</sub>	37.00%
		Malic Acid	37.00%
		EXACT 4023	26.00%
		CaCO <sub>3</sub>	35.00%
		Malic Acid	38.90%
		PolyOx 750	5.00%
		EXACT 4023	21.10%
		CaCO <sub>3</sub>	35.00%

		Malic Acid	38.90%
		PolyOx 750	5.00%
		PP 3505	21.10%
		CaCO <sub>3</sub>	35.00%
		Citric Acid	37.10%
		PolyOx 750	5.00%
		EXACT 4023	22.90%
		CaCO <sub>3</sub>	35.00%
		Citric Acid	37.10%
		PolyOx 750	5.00%
		PP 3505	22.90%
		CaCO <sub>3</sub>	35.00%
		Malic Acid	38.90%
		PP 3505	26.10%
		CaCO <sub>3</sub>	35.00%
		Citric Acid	37.10%
		PP 3505	27.90%
		CaCO <sub>3</sub>	37.00%
		Citric Acid	31.00%
		PolyOx 750	5.00%
		Calcium Hydroxide (anhydrous)	6.00%
		EXACT 4023	21.10%

### **Example 1**

In one embodiment, a suitable composition for a three-phase polymer film is set forth below. All compositions are set forth on a mass/mass basis.

<u>Type (component "A")</u>	<u>Manufacturer</u>	<u>Product Number</u>
27% Ethylene-Olefin Copolymer	Exxon Chemical Company	4023

### **Type (component "B")**

3% Ethylene oxide	Union Carbide Corporation	WSRN750
30%Acrylic Polymer	BF Goodrich	934PNF

### **Type (component "C")**

40% CaCO <sub>3</sub>	Sigma-Aldrich Chemical Company	202932
-----------------------	--------------------------------	--------

The calcium carbonate is included in the composition in particulate form. All sample components are solid phase. The above-mentioned formulation was hand mixed. The processing of formulation was done using a single screw extruder. The temperature profile across the extruder was maintained at 250°F. The single screw extruder was primed/purged with Exact 4023. The formulation was then added to the extruder through the hopper. About 5 grams of extrudate was then collected coming out of the die.

The collected extrudate was placed in between the two platen of the compression press and it was gently squeezed at 5000 psi and for about 10 seconds. Thickness of the sample was maintained by means of shims. The sample was then quenched. The final circular shape weighing 1 gram having an inch diameter and 0.065-inch thickness was cut by means of circular stamp. The circular shapes can be fixed to the cap liner in the way previously discussed.

## **Example 2**

### **Sample Vials**

A total of 90 vials were used for the CO<sub>2</sub> release study. 55 vials, used as controls, were purged with nitrogen. The other 35 vials (replicate for each time frame) were set up for CO<sub>2</sub> release test. Vial volume 120cc. The vials are made of glass. The vials had crimped aluminum seals with red rubber septa in it.

### **Test conditions**

Temperature in the room in which testing was done was maintained between 22-23°C. Each vials had half a piece of round 90 mm diameter filter paper, moist with 500μL of water. Each vial had sample inside it. The lid of the vial was crimped tight after putting the sample and the filter paper inside it. The test was run over a period of 4 weeks.

**Sample weight-** The samples were composed of the composition of Example 1. All samples of the CO<sub>2</sub> releasing composition were weighed prior to the placement into the vial. Approximately 0.5 to 1.0 grams of sample was used for the experiment.

### **CO<sub>2</sub> release measurement**

A MOCON PAC CHECK-650 Dual Head Space Analyzer device was used to measure the amount of carbon dioxide released by the sample. This analyzer employs an IR source and IR detector in the determination of the amount of CO<sub>2</sub> present at the time the sample is inspected.

The % carbon dioxide released by the sample was measured after 1 day, 2 days, 3 days, 5days, 7 days, 14 days, 28 days duration. The 55 control vial samples (5 replicates for each time period) are checked at 1 hr, 2hr, 4 hr, 8 hr, 24 hr, 2 days, 4 days, 7 days, 14 days, 21 days and 28 days for percent of oxygen inside the vials or leakage through the vial.

Samples were removed from the vial and inserted into the MOCON instrument to detect the amount of CO<sub>2</sub> present in the following manner. A 10 cc gas tight syringe was used. The syringe was emptied by pushing the plunger all the way in. The syringe needle was inserted through the rubber septa located in the lid of the aluminum-crimped seal. The syringe plunger was slowly pulled out until a gas sample of 5cc or more is in the syringe. The needle was inserted into the injection filter assembly on the front panel of the PAC CHECK 650 until the needle will not go any further. The syringe plunger was pushed to inject the sample into the PAC CHECK 650 at a rate of approximately 1cc/second. As the sample is injected, the readouts on the MOCON instrument for the CO<sub>2</sub> and O<sub>2</sub> display are checked to insure the device is operating properly. These indicate that the sample was properly injected into the PAC CHECK 650. The Concentration of Oxygen and Carbon dioxide is displayed in the front panel in %.

The control sample vials were filled with nitrogen. These vials are checked at above-mentioned time for percent of oxygen inside the vials or leakage through the vial. The % carbon dioxide in the room atmosphere was measured at the beginning of the experiment. The % carbon dioxide released by the sample was measured at selected intervals. The volume of the air inside the vial was calculated by filling the vial with water and weighing it after tarring the vial weight with filter paper and sample inside it. The difference in the % CO<sub>2</sub> reading was converted to the volume of the CO<sub>2</sub> released based on the vial volume. Then based on the weight of the sample the equivalent volume of CO<sub>2</sub> released/gm of the sample was calculated after implying a correction factor for the increase in pressure inside the vial.

**Results** The result of the leakage test for control samples is shown in Figure 2.

The data collected for the carbon dioxide release is summarized in Table 1.

The Figure 1 represents the results for CO<sub>2</sub> release.

The data show that the amount of CO<sub>2</sub> released/gm of the sample is-32.27 cc based on 14 days data.

#### **Findings-**

The data show that the CSP CO<sub>2</sub> releaser film has the capacity of:

**Table1**

Time (Days)	CC of CO <sub>2</sub> released/gm of sample
1	8.74
2	13.60
3	16.51
5	25.04

7	29.15
14	32.27
28	Due

### **Example 3**

The following are various CO<sub>2</sub> releasing compositions illustrating the present invention. The table also shows the releasing profiles with various acids. Figure 3 is a graph showing these releasing profiles.

1/15/01	CO2 release in 48 hours by actives		
Ratio of NaHCO <sub>3</sub> :acid=5:1	(with moist filter paper)	cc of CO2 released/gm of active <b>NAHCO<sub>3</sub></b>	cc of CO2 released/gm of active <b>CaCO<sub>3</sub> (larger particle size)</b>
citric acid	24 hours	70.05	34.28
	48 hours	<b>83.22</b>	<b>43.23</b>
maleic acid	24 hours	56.6	52.83
	48 hours	<b>65.01</b>	<b>39.12</b>
malic acid	24 hours	78.23	39.64
	48 hours	<b>75.61</b>	<b>33.64</b>
fumaric acid	24 hours	55.94	7.14
	48 hours	<b>81.23</b>	<b>10.05</b>
carbopol(polyacrylic acid)	24 hours	30.09	12.5
	48 hours	<b>39.42</b>	<b>4.59</b>